PROCEEDINGS OF THE AMERICAN CHEMICAL SOCIETY.

Regular Meeting, February 2, 1883.

Mr. James H. Stebbins, Jr., in chair.

The minutes of November, December and January meetings were read and approved.

The following gentlemen were unanimously elected: Messrs. Percy Neymann, J. Howard Wainwright, Max Schwartz, Vincent M. Picabia, Jas. L. Howe, Andrew Peters, F. T. Walsh and T. Lambert.

Mr. Chas. Eimer (130 E. 18th St.) was then nominated by James H. Stebbins, Jr., Wm. Rupp, and Aug. Eimer.

Remarks upon the last paper were then made by Messrs. Waller, Casamajor and Gladding.

Mr. Gladding explained at length his plan of sampling soap and determining water, as follows:

To prepare an average sample, one inch in length is taken from the middle of the bar and is thoroughly mixed with a stout spatula.

For the determination of the water, a small beaker (100 cc. capacity) is accurately weighed. Into this is then weighed a half inch in depth of pure dry quartz sand, and also a small glass rod. About 5 grams of the thoroughly mixed sample is then placed in the beaker and accurately weighed. 25 cc. of 95 per cent. alcohol are then poured into the beaker and the whole placed in a water oven. The alcohol dissolves the soap, and the contents of the beaker are stirred at intervals by means of the glass rod. As the alcohol evaporates, the soap is left in a very thin coat upon the particles of sand and the expulsion of water is very complete. The perpendicular sides of the beaker prevent any loss of alcohol from creeping.

Mr. Gladding suggested that the residue might then be treated with hot petroleum ether, the ether decanted and the treatment repeated several times and the residue again dried. The loss would be uncombined fat. The residue might then be dissolved in 95 per cent. alcohol, the solution filtered and the filtrate and residue carried through the several stages proposed by Dr. Leeds.

As he had had trouble determining the uncombined alkali by the titration recommended by Dr. Leeds, the precipitation of this by a

current of CO_2, followed by filtration and estimation of the precipitated carbonates, might be found more reliable.

The following papers were read :

1. "On the Action of Diazoanisole Chloride upon Phenoles and their Substitution Products," by Jas. H. Stebbins, Jr.
2. "On Acrolein Urea, with remarks upon Prof. Hugo Schiff's publications upon Condensed Ureas," by Dr. A. R. Leeds.
3. "Detection of Anhydrous Glucose mixed with Refined Cane Sugar," by P. Casamajor.
4. "A new Scheme of Soap Analysis, with a preliminary discussion of former methods," by Dr. A. R. Leeds.

The resignation of Mr. M. Benjamin was accepted.

The meeting then adjourned.

THOS. S. GLADDING,
Recording Secretary.

ACTION OF DIAZOANISOLE CHLORIDE UPON PHENOLES, AND THEIR SUBSTITUTION PRODUCTS.

BY JAMES H. STEBBINS, JR.

On the 29th day of Jan., 1879, Peter Griess filed a specification for the production of a red coloring matter, obtained by the action of diazoanisole chloride upon an alkaline solution of the beta sulpho acid of beta naphthole.

On Feb. 12, 1879, he filed another specification, for the production of a crimson dyestuff, formed by the action of diazoanisole chloride upon an alkaline solution of the disulpho acid of beta naphthole.

These two specifications, therefore, teach us that the diazo-ethers of phenoles can react upon phenoles, as well as the more ordinary diazo-compounds.

Since the dyestuffs produced in the above manner are but few in number, I thought it worth my while to complete, if possible, the series already so well begun.

For this purpose the following experiments were made :

Orcine-azo-benzole-di-soda salt.

12.3 grms. anisidine were dissolved in 17.9 grms. HCl. (1.20 sp. gr.), diluted with 200 cc. water.